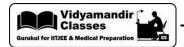


Date Planned : / /	Daily Tutorial Sheet-3	Expected Duration : 30 Min		
Actual Date of Attempt : / /	Level-1	Exact Duration :		

31.	The rea	te reaction of sodium is highly exothermic with water. The rate of reaction is lowered by :							
	(A)	lowering the temperature	(B)	mixing with alcohol					
	(C)	mixing with acetic acid	(D)	making an amalgam					
32 .	RbO ₂	is:			\odot				
	(A)	Superoxide and paramagnetic	(B)	Superoxide and diamagnetic					
	(C)	Peroxide and paramagnetic	(D)	Peroxide and diamagnetic					
33.	KOH is	preferably used to absorb ${\rm CO}_2$ gas beca	use :		\odot				
	(A)	KHCO_3 is soluble in water while NaHCO_3 is sparingly soluble in water							
	(B)	KOH is cheaper than NaOH							
	(C) KOH is stronger base than NaOH								
	(D)	KOH is more soluble than NaOH in water	er						
34.	Washing soda on heating:								
	(A)	Releases CO gas	(B)	Releases CO ₂ gas					
	(C)	Releases water vapour	(D)	Both (A) and (B)					
35.	When s	sodium hydroxide solution is electrolysed	:		\odot				
	(A)	hydrogen is discharged at cathode	(B)	hydrogen is discharged at anode					
	(C)	sodium is liberated at anode	(D)	hydrogen is not liberated					
36.	${\rm LiAlH_4}$	is used as:			\odot				
	(A)	an oxidizing agent	(B)	a reducing agent					
	(C)	a mordant	(D)	a water softener					
37.	Sodium	n carbonate can be manufactured by Solve:	vay proc	ess but potassium carbonate cannot be	prepared				
	(A)	${ m K_2CO_3}$ is more soluble	(B)	${ m K_2CO_3}$ is less soluble					
	(C)	${\rm KHCO_3}$ is more soluble than ${\rm NaHCO_3}$	(D)	${\rm KHCO_3}$ is less soluble than ${\rm NaHCO_3}$					
38.	Which	of the following increases in magnitude a	s the ato	omic number of alkali metals increases?					
	(A)	Electronegativity	(B)	First ionization potential					
	(C)	Ionic radius	(D)	Melting point					
39.	Sodium	n has as compared to potassium:			\odot				
	(A)	less electronegativity	(B)	More ionization potential					
	(C)	larger atomic radius	(D)	lower melting point					
40 .	The metallic lustre exhibited by sodium is explained by :								
	(A)	diffusion of sodium ions	(B)	oscillation of mobile valence electrons					
	(C)	existence of free protons	(D)	existence of body centred cubic lattice					



41.	In the electrolysis of NaCl solution for the manufacture of NaOH, the ion discharged at graphite cathode									
	is:									
	(A)	Na ⁺	(B)	Cl ⁻	(C)	H^+	(D)	O^{2-}		
42 .	Among LiCl, RbCl, $BeCl_2$ and $MgCl_2$ the compounds with greatest and least ionic character respective								spectively	
	are :								\odot	
	(A)	LiCl, RbCl	(B)	${\rm RbCl~BeCl}_2,$	(C)	RbCl, MgCl_2	(D)	$MgCl_2,BeCl_2$		
43 .	Stable oxide is obtained by heating the carbonate of the elements:							\odot		
	(A)	Li	(B)	Na	(C)	K	(D)	Rb		
44.	Which reacts directly with nitrogen to form nitride?							\odot		
	(A)	Na	(B)	Li	(C)	K	(D)	Rb		
45 .	Which o	of the following h	as the h	ighest melting p	oint?					
	(A)	NaCl	(B)	NaF	(C)	NaBr	(D)	Nal		